



QUIZZES

Practice Test-1(Liquids)



10 Questions



7 min

Topics

Properties of liquids, Intermolecular forces (Van DER WAAL's Forces)

Start Quiz

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06 : 58



1/10



7 min



Hint

Q : London dispersion forces are strongest in



F_2



Cl_2



Br_2



I_2

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06 : 57



2/10



7 min



Hint

Q : Which of the followings does not match

A

H_2O and Na^+ \longrightarrow Ion dipole force

B

CH_3COCH_3 and CH_3COCH_3 \longrightarrow dipole-dipole force

C

HCl and Ar \longrightarrow Dipole-dipole force

D

$\text{C}_6\text{H}_{12}\text{O}_6$ and H_2O \longrightarrow Hydrogen bonding

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06 : 55



3/10



7 min



Hint

Q : A substance that is soluble in water due to dipole-induced dipole force



$C_6H_{12}O_6$



HCl



Kr



H_2S

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06 : 54



4/10



7 min



Hint

Q : The polarizabilities of elements mostly increase down the group due to the



Increase in the atomic number



Increase in the number of shell



Increase in the number of protons



The behaviour of element remains the same

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06 : 53



5/10



7 min



Hint

Q : Which of the following inter molecular forces may be present in molecular solids



Dipole – dipole forces



Van der Waal's forces



H- bonding



All of the these

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06 : 51



6/10



7 min



Hint

Q : Which liquid will boil at lower temperature



Benzene



Ethanol



Acetic acid



Acetone

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06 : 50



7/10



7 min



Hint

Q : Molar heat of vapourization is the highest for



Chlorine



Bromine



Iodine



Fluorine

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06 : 49



8/10



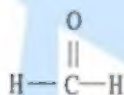
7 min



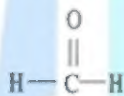
Hint

Q : Which one of the following substances will have hydrogen bonding as one of its intermolecular forces among itself

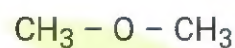
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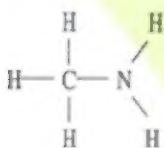
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06 : 47



9/10



7 min



Hint

Q : The inter-molecular attractive forces vary in the order



Water < alcohol < ether



Alcohol < water < ether



Ether < alcohol < water



Ether < water < Alcohol

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06 : 44



10/10



7 min



Hint

Q : Which property of liquids leads to the development of thermometers



Pressure variation of liquids with temperature



Liquids are visible



Expansion of liquid at high temperature



Liquid can not be compressed

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QUIZ RESULT

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10



7 min



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0 sec



0/10



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Result Detail

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Correct

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Incorrect

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Unattempted

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← Practice Test-1(Liquids)



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Incorrect



1/10

Q : London dispersion forces are strongest in



F₂



Cl₂



Br₂



I₂

Explanation

London dispersion forces are strongest in I₂.

London dispersion forces  polarizability

I₂ is more polarizable due to large size of molecule so it is solid at room temperature.

Intermolecular forces or polarizability order

I₂ > Br₂ > Cl₂ > F₂

← Practice Test-1(Liquids)



Correct



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Incorrect



2/10

Q : Which of the followings does not match



H_2O and Na^+ \longrightarrow Ion dipole force



CH_3COCH_3 and CH_3COCH_3 \longrightarrow dipole-dipole force



HCl and Ar \longrightarrow Dipole-dipole force



$\text{C}_6\text{H}_{12}\text{O}_6$ and H_2O \longrightarrow Hydrogen bonding

Explanation

HCl and Ar \longrightarrow Dipole - induced dipole force

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3/10

Q : A substance that is soluble in water due to dipole-induced dipole force



$C_6H_{12}O_6$



HCl



Kr



H_2S

Explanation

Krypton is non-polar so soluble in water due to dipole – induce dipole forces.

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4/10

Q : The polarizabilities of elements mostly increase down the group due to the



Increase in the atomic number



Increase in the number of shell



Increase in the number of protons



The behaviour of element remains the same

Explanation

Polarizability is directly related to size of electronic cloud

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5/10

Q : Which of the following inter molecular forces may be present in molecular solids



Dipole – dipole forces



Van der Waal's forces



H- bonding



All of the these

Explanation

Types of molecules polar and non-polar molecules solid and all type of forces present between molecules solid.

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Unattempted



Incorrect



6/10

Q : Which liquid will boil at lower temperature



Benzene



Ethanol



Acetic acid



Acetone

Explanation

Acetone has weaker forces of attraction among all of given compounds

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Incorrect



7/10

Q : Molar heat of vapourization is the highest for



Chlorine



Bromine



Iodine



Fluorine

Explanation

Greater the size of molecule, greater the polarizability, greater the forces of attractions hence greater the heat of vapourization

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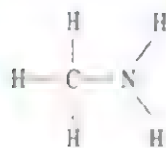
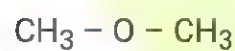
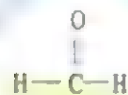
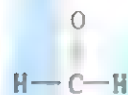


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8/10

Q : Which one of the following substances will have hydrogen bonding as one of its intermolecular forces among itself



Explanation

$\text{CH}_3\text{-NH}_2$ have hydrogen bonding

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Correct



Unattempted



Incorrect



9/10

Q : The inter-molecular attractive forces vary in the order



Water < alcohol < ether



Alcohol < water < ether



Ether < alcohol < water



Ether < water < Alcohol

Explanation

Water > alcohol > ether is the correct order of force of attraction, water and alcohol have hydrogen bonding while ether has London forces.

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← Practice Test-1(Liquids)



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Incorrect



10/10

Q : Which property of liquids leads to the development of thermometers



Pressure variation of liquids with temperature



Liquids are visible



Expansion of liquid at high temperature



Liquid can not be compressed

Explanation

development of thermometers is due to expansion of liquid at high temperature

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Practice Test-2(Liquids)



10 Questions



7 min

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06 : 58



1/10



7 min



Hint

Q : Heat of vaporization is minimum for



HF



HCl



HBr



HI

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06 : 56



2/10



7 min



Hint

Q : If we provide very large amount of heat to a liquid, its vapor pressure at its boiling point



Remains the same



Increases



Decreases



Varies abnormally

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06 : 55



3/10



7 min



Hint

Q : Vapour pressure of liquid is measured when liquid and the vapours are in equilibrium it means that



Liquid and vapours have same value of kinetic energy



Liquid and vapours have same heat content



Rate of evaporation is equal to the rate of condensation



Rate of evaporation and condensation are different

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06 : 53



4/10



7 min



Hint

Q : When a liquid is evaporated



Temperature of liquid increases



Temperature of liquid decreases



Liquid molecules becomes more energetic



Both "A" and "C"

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06 : 51



5/10



7 min



Hint

Q : The boiling point of water would be highest at



Murree



Gawadar



Mount everest



Siachin

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06 : 50



6/10



7 min



Hint

Q : Which of the following compound has higher boiling point



H₂O



H₂Se



H₂S



CH₄

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06 : 48



7/10



7 min



Hint

Q : The lowest vapour pressure is of



H₂O



HCl



CH₃OCH₃



CH₂-CH₂
OH OH

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06 : 47



8/10



7 min



Hint

Q : Hydrogen bonding links one spiral (DNA) to the other. The H-bonding is more dominant between



C and H



N and H



O and H



O and N

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06 : 45



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7 min



Hint

Q : Acetic acid exists as dimer in benzene due to



A Condensation reaction



B Hydrogen bonding



C Presence of carboxyl group



D Presence of hydrogen atom at α - carbon

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06 : 42



10/10



7 min



Hint

Q : Water has maximum density at



0°C



- 4°C



100°C



4°C

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QUIZ RESULT

Practice Test-2(Liquids)



10



7 min



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Result Detail

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Incorrect

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← Practice Test-2(Liquids)



Correct



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Incorrect



1/10

Q : Heat of vaporization is minimum for



HF



HCl



HBr



HI

Explanation

HCl have weakest forces which makes HCl most volatile and have lowest heat of vaporization

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



2/10

Q : If we provide very large amount of heat to a liquid, its vapor pressure at its boiling point



Remains the same



Increases



Decreases



Varies abnormally

Explanation

Boiling point is depend upon external pressure and intermolecular force

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← Practice Test-2(Liquids)



Correct



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Incorrect



3/10

Q : Vapour pressure of liquid is measured when liquid and the vapours are in equilibrium it means that



Liquid and vapours have same value of kinetic energy



Liquid and vapours have same heat content



Rate of evaporation is equal to the rate of condensation



Rate of evaporation and condensation are different

Explanation

Evaporation and condensation are reversible.

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



4/10

Q : When a liquid is evaporated



Temperature of liquid increases



Temperature of liquid decreases



Liquid molecules becomes more energetic



Both "A" and "C"

Explanation

Evaporation is endothermic.

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



5/10

Q : The boiling point of water would be highest at



Murree



Gawadar



Mount everest



Siachin

Explanation

Because of high external pressure at Gawadar

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← Practice Test-2(Liquids)



Correct



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Incorrect



6/10

Q : Which of the following compound has higher boiling point



H₂O



H₂Se



H₂S



CH₄

Explanation

Water has hydrogen bonding

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



7/10

Q : The lowest vapour pressure is of



H₂O



HCl



CH₃OCH₃



CH₂-CH₂
OH OH

Explanation

Greater the forces of attraction lower the vapour pressure

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



8/10

Q : Hydrogen bonding links one spiral (DNA) to the other. The H-bonding is more dominant between



C and H



N and H



O and H



O and N

Explanation

The H-bonding is more dominant between O and H.
Hydrogen bonding is also present between nitrogen and hydrogen.

← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



9/10

Q : Acetic acid exists as dimer in benzene due to



Condensation reaction



Hydrogen bonding



Presence of carboxyl group



Presence of hydrogen atom at α - carbon

Explanation

Acetic acid form hydrogen bond with its other molecule in non-polar solvent

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



10/10

Q : Water has maximum density at



0°C



- 4°C



100°C



4°C

Explanation

Below 4°C empty spaces are created. Above and below 4°C molecules of H₂O are at larger distances as compared to distance at 4°C

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Practice Test-1(Solids)



10 Questions



7 min

Topics

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06 : 56



1/10



7 min



Hint

Q : Which is a pseudo solid?



Rock salt



Aluminum nitride



Glue



Graphite

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06 : 55



2/10



7 min



Hint

Q : Glass is called super cooled liquid. The reason is that glass has



Definite volume



Definite shape



Crystalline structure



No crystalline structure

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06 : 53



3/10



7 min



Hint

Q :

Crystalline solids have different properties

I. Allotropy
Polymorphism

II. Isomorphism

III.

Ionic compounds show which of the above properties



I, II



II, III



I, II, III



I, III

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06 : 51



4/10



7 min



Hint

Q : Which is not a property of isomorphous solids?



Have same atomic ratio



same crystalline shape



Same chemical property



same geometry of anion

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06 : 48



5/10



7 min



Hint

Q : _____ may be isotropic under ordinary conditions



Diamond



Borax



Bromine



Sodium Chloride

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06 : 46



6/10



7 min



Hint

Q : Which property is incorrect about silica



High transparency to light



Very low thermal expansion



Excellent insulator



Soluble in water

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06 : 45



7/10



7 min



Hint

Q : Diamond and graphite are



Isomorphous



Polymorphous



Allotropes



Both "B" and "C"

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06 : 43



8/10



7 min



Hint

Q : In a crystal lattice the correct sequence of bond angles is



$bc = \alpha, ca = \beta, ab = \gamma$



$bc = \beta, ca = \alpha, ab = \gamma$



$bc = \gamma, ca = \beta, ab = \alpha$



All are possible

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06 : 41



9/10



7 min



Hint

Q : The arrangement of particles as a points in a crystal is called



Unit cell



Crystal lattice



Space lattice



Both B and C

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06 : 38



10/10



7 min



Hint

Q : Unit cell is the smallest part, has all features of the entire



Molecule



Compound



Atom



Crystal

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QUIZ RESULT

Practice Test-1(Solids)



10



7 min



11-Apr-2021



0 sec



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Result Detail

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Correct

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Incorrect

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Unattempted

10

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Chemistry

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← Practice Test-1(Solids)



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1/10

Q : Which is a pseudo solid?



Rock salt



Aluminum nitride



Glue



Graphite

Explanation

Glue is amorphous solid

Pseudo solid is also called amorphous solid

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2/10

Q : Glass is called super cooled liquid. The reason is that glass has



Definite volume



Definite shape



Crystalline structure



No crystalline structure

Explanation

Glass is solidified by the process annealing, due to which it has non crystalline structure

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← Practice Test-1(Solids)



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3/10

Q :

Crystalline solids have different properties

I. Allotropy
Polymorphism

II. Isomorphism

III.

Ionic compounds show which of the above properties



I, II



II, III



I, II, III



I, III

Explanation

Ionic solids show isomorphism and polymorphism.
Allotropy shown by element.



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Q : Which is not a property of isomorphous solids?



Have same atomic ratio



same crystalline shape



Same chemical property



same geometry of anion

Explanation

Isomorphous solids generally have same atomic ratio and crystalline shape but they are chemically different substances e.g. NaNO_3 and KNO_3 are isomorphs of each other. Isomorphous solids do not have same chemical properties.

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Q : _____ may be isotropic under ordinary conditions



Diamond



Borax



Bromine



Sodium Chloride

Explanation

Bromine exist as liquid at normal temperature so it may be isotropic

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Q : Which property is incorrect about silica



High transparency to light



Very low thermal expansion



Excellent insulator



Soluble in water

Explanation

Silica is giant structure and insoluble in water.

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7/10

Q : Diamond and graphite are



Isomorphous



Polymorphous



Allotropes



Both "B" and "C"

Explanation

Diamond graphite both consist of C-atoms.

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Q : In a crystal lattice the correct sequence of bond angles is



$bc = \alpha, ca = \beta, ab = \gamma$



$bc = \beta, ca = \alpha, ab = \gamma$



$bc = \gamma, ca = \beta, ab = \alpha$



All are possible

Explanation

α -angle lies between side b & c

β -angle lies between side a & c

γ -angle lies between side b & a

← Practice Test-1(Solids)



Correct



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Incorrect



9/10

Q : The arrangement of particles as a points in a crystal is called



Unit cell



Crystal lattice



Space lattice



Both B and C

Explanation

The arrangement of particles as a points in a crystal is called Crystal lattice or Space lattice

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10/10

Q : Unit cell is the smallest part, has all features of the entire



Molecule



Compound



Atom



Crystal

Explanation

Unit cell is the smallest part, has all features of the entire crystal

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QUIZZES

Practice Test-2(Liquids)



10 Questions



7 min

Topics

Classification

Start Quiz

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06 : 58



1/10



7 min



Hint

Q : The crystal system for $a \neq b \neq c$ and $\alpha \neq \beta \neq \gamma \neq 90^\circ$



Cubic



Orthorhombic



Tetragonal



Triclinic

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06 : 56



2/10



7 min



Hint

Q : Which one of the following set of substances shows hexagonal close packing structure



Li, Na, K, Rb



Zn, Cd, Mg



Ag, Cu, Au, Al, Fe



NaCl

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06 : 55



3/10



7 min



Hint

Q : Which of the following pair of solids are giant covalent solids?



Zn, Cu



CaO, CsF



AlN, SiC



C₁₀H₈, N₂

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06 : 53



4/10



7 min



Hint

Q : Which one of the following may not be related to ionic compounds



Transition temperature



Isomorphison



Polymorphism



Isomerism

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06 : 51



5/10



7 min



Hint

Q : Ionic solids do not conduct the electrical current because



A Ions don't have translatory motion



B Free electrons are less



C The coordination number of ion is very high



D Strong covalent bonds are present in their structure

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06 : 49



6/10



7 min



Hint

Q : Molecular crystals are



Soft



Hard



Highly dense



Closely packed

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06 : 48



7/10



7 min



Hint

Q : The crystals formed due to London's forces are



Ionic



Covalent



Molecular



Metallic

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06 : 46



8/10



7 min



Hint

Q : Crystalline solids are classified on the basis of bonding into



Two types



Seven types



Five types



Four types

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06 : 44



9/10



7 min



Hint

Q : The molecules of CO_2 in dry ice form the



Ionic crystals



Covalent crystals



Molecular crystals



Any type of crystals

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06 : 42



10/10



7 min



Hint

Q : Diamond, Silicon carbide are insoluble in most of the solvents because



They have very high lattice energy



They have very big size



They do not interact with the solvent



All of these

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QUIZ RESULT

Practice Test-2(Liquids)



10



7 min



11-Apr-2021



0 sec



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Result Detail

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Correct

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Incorrect

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Unattempted

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Chemistry

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



1/10

Q: The crystal system for $a \neq b \neq c$ and $\alpha \neq \beta \neq \gamma \neq 90^\circ$



Cubic



Orthorhombic



Tetragonal



Triclinic

Explanation

In triclinic system all angles and all lengths are not equal

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← Practice Test-2(Liquids)



Correct



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Incorrect



2/10

Q : Which one of the following set of substances shows hexagonal close packing structure



Li, Na, K, Rb



Zn, Cd, Mg



Ag, Cu, Au, Al, Fe



NaCl

Explanation

Zn, Cd, Mg is the set of substances shows hexagonal close packing structure.

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← Practice Test-2(Liquids)



Correct



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Incorrect



3/10

Q : Which of the following pair of solids are giant covalent solids?



Zn, Cu



CaO, CsF



AlN, SiC



C₁₀H₈, N₂

Explanation

AlN, SiC are giant covalent solids.

Zn, Cu are metallic solids

CaO, CsF are ionic solids

C₁₀H₈ is molecular solids

N₂ is gas at room temperature

← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



4/10

Q : Which one of the following may not be related to ionic compounds



Transition temperature



Isomorphison



Polymorphism



Isomerism

Explanation

In ionic compound there is non-directional forces so did not show isomerism

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← Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



5/10

Q : Ionic solids do not conduct the electrical current because



Ions don't have translatory motion



Free electrons are less



The coordination number of ion is very high



Strong covalent bonds are present in their structure

Explanation

Ionic bond is very strong so, it does not have translatory motion

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Practice Test-2(Liquids)



Correct



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Incorrect



6/10

Q : Molecular crystals are

A

Soft

B

Hard

C

Highly dense

D

Closely packed

Explanation

Molecular solids have very weak intermolecular forces due to which molecular solids are soft

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Practice Test-2(Liquids)



Correct



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Incorrect



8/10

Q : Crystalline solids are classified on the basis of bonding into

A

Two types

B

Seven types

C

Five types

D

Four types

Explanation

Ionic bond

Covalent solid

Molecular solid

Metallic solid



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Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



9/10

Q : The molecules of CO_2 in dry ice form the

A

Ionic crystals

B

Covalent crystals

C

Molecular crystals

D

Any type of crystals

Explanation

Solidified CO_2 is non polar molecule and has L.D.F.

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Practice Test-2(Liquids)



Correct



Unattempted



Incorrect



10/10

Q : Diamond, Silicon carbide are insoluble in most of the solvents because

A

They have very high lattice energy

B

They have very big size

C

They do not interact with the solvent

D

All of these

Explanation

Non-polar and Macromolecules having giant structure.

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